Lecture 6: Linear Algebra I

Reading:
Kreyszig Sections: 4.0, 7.1, 7.2, 7.3, 7.4, 7.5

Vectors

Vectors as a list of associated information

\[ \vec{x} = \left( \begin{array}{c} \text{number of steps to the east} \\ \text{number of steps to the north} \\ \text{number steps up vertical ladder} \end{array} \right) \]  

(6-1)

\[ \vec{x} = \left( \begin{array}{c} 3 \\ 2.4 \\ 1.5 \end{array} \right) \]
determines position

(6-2)

The vector above is just one example of a position vector. We could also use coordinate systems that differ from the Cartesian \((x, y, z)\) to represent the location. For example, the location in a cylindrical coordinate system could be written as

\[ \vec{x} = \left( \begin{array}{c} x \\ y \\ z \end{array} \right) = \left( \begin{array}{c} r \cos \theta \\ r \sin \theta \\ z \end{array} \right) \]  

(6-3)

i.e., as a Cartesian vector in terms of the cylindrical coordinates \((r, \theta, z)\).

The position could also be written as a cylindrical, or polar vector

\[ \vec{x} = \left( \begin{array}{c} r \\ \theta \\ z \end{array} \right) = \left( \begin{array}{c} \sqrt{x^2 + y^2} \\ \tan^{-1} \frac{y}{x} \\ z \end{array} \right) \]  

(6-4)

where the last term is the polar vector in terms of the Cartesian coordinates. Similar rules would apply for other coordinate systems like spherical, elliptic, etc.
However, vectors need not represent position at all, for example:

$$\vec{n} = \begin{pmatrix}
\text{number of Hydrogen atoms} \\
\text{number of Helium atoms} \\
\text{number of Lithium atoms} \\
\vdots \\
\text{number of Plutonium atoms}
\end{pmatrix}$$  \hspace{1cm} (6-5)

Scalar multiplication

$$\frac{1}{\text{Navag.}} \vec{n} \equiv \begin{pmatrix}
\frac{\text{number of H}}{\text{Navag.}} \\
\frac{\text{number of He}}{\text{Navag.}} \\
\frac{\text{number of Li}}{\text{Navag.}} \\
\vdots \\
\frac{\text{number of Pu}}{\text{Navag.}}
\end{pmatrix} = \begin{pmatrix}
\text{moles of H} \\
\text{moles of He} \\
\text{moles of Li} \\
\vdots \\
\text{moles of Pu}
\end{pmatrix} = \vec{m} \hspace{1cm} (6-6)$$

Vector norms

$$\|\vec{x}\| \equiv (x_1^2 + x_2^2 + \ldots + x_k^2)^{1/2} = \text{euclidean separation (also called l2-norm)} \hspace{1cm} (6-7)$$

$$\|\vec{n}\| \equiv n_{\text{H}} + n_{\text{He}} + \ldots + n_{\text{Pu}} = \text{total number of atoms (related to the Manhattan norm)} \hspace{1cm} (6-8)$$

Unit vectors
unit direction vector

\[
\hat{x} = \frac{\vec{x}}{\| \vec{x} \|}
\]
mole fraction composition

\[
\hat{X} = \frac{\vec{N}}{\| \vec{N} \|}
\]

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**Extra Information and Notes**

*Potentially interesting but currently unnecessary*

If \( \mathbb{R} \) stands for the set of all real numbers (i.e., 0, -1.6, \( \pi/2 \), etc.), then we can use a shorthand to specify the position vector, \( \vec{x} \in \mathbb{R}^N \) (e.g., each of the \( N \) entries in the vector of length \( N \) must be a real number, or must be in the set of real numbers, \( \| \vec{x} \| \in \mathbb{R} \)).

For the unit (direction) vector: \( \hat{x} = \{ \vec{x} \in \mathbb{R}^3 \mid \| \vec{x} \| = 1 \} \) (i.e., the unit direction vector is the set of all position vectors such that their length is unity—or, the unit direction vector is the subset of all position vectors that lie on the unit sphere. \( \vec{x} \) and \( \hat{x} \) have the same number of entries, but compared to \( \vec{x} \), the number of independent entries in \( \hat{x} \) is smaller by one.

For the case of the composition vector, it is unphysical to have a negative number of atoms, therefore the mole fraction vector \( \vec{n} \in (\mathbb{R}^+) \) elements (\( \mathbb{R}^+ \) is the real non-negative numbers) and \( \vec{m} \in (\mathbb{R}^+)\) elements-1.

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### Matrices and Matrix Operations

Consider methane (CH\(_4\)), propane (C\(_3\)H\(_8\)), and butane (C\(_4\)H\(_{10}\)).

\[
M_{HC} = \begin{pmatrix}
\text{H-column} & \text{C-column}
\end{pmatrix}
\begin{pmatrix}
\text{number of H} & \text{number of C}
\text{methane molecule} & \text{methane molecule}
\text{propane molecule} & \text{propane molecule}
\text{butane molecule} & \text{butane molecule}
\end{pmatrix}
\begin{pmatrix}
\text{methane row} \\
\text{propane row} \\
\text{butane row}
\end{pmatrix}
\]

\[
M_{HC} = \begin{pmatrix} 4 & 8 & 10 \\ 1 & 3 & 4 \end{pmatrix} = \begin{pmatrix} M_{11} & M_{12} \\ M_{21} & M_{22} \\ M_{31} & M_{32} \end{pmatrix}
\]
Matrices as a linear transformation of a vector

\[ \hat{\vec{N}}_{HC} = ( \text{number of methanes}, \text{number of propanes}, \text{number of butanes}) \]  
\[ = (N_{HC\ m}, N_{HC\ p}, N_{HC\ b}) \]  
\[ = (N_{HC\ 1}, N_{HC\ 2}, N_{HC\ 3}) \]  

(6-13) (6-14) (6-15) (6-16)

\[ \hat{\vec{N}}_{HC} M_{HC} \equiv \sum_{i=1}^{3} N_{HC\ i} M_{HC\ ij} = \vec{N} \]  

(6-17)

The “summation” convention is often used, where a repeated index is summed over all its possible values:

\[ \sum_{i=1}^{p} N_{HC\ i} M_{HC\ ij} \equiv N_{HC\ i} M_{HC\ ij} = N_j \]  

(6-18)

For example, suppose

\[ \hat{\vec{N}}_{HC} = (1.2 \times 10^{12} \text{ molecules methane}, 2.3 \times 10^{13} \text{ molecules propane}, 3.4 \times 10^{14} \text{ molecules butane}) \]  

(6-19)
\[ N_{HC}^{\top} M_{HC} = \]
\[
\left( 1.2 \times 10^{14} \text{ methanes, } 2.3 \times 10^{13} \text{ propanes, } 3.4 \times 10^{12} \text{ butanes} \right) \left( \begin{array}{cc} 4 \text{ atoms H} & 1 \text{ atom C} \\ \text{methane} & \text{methane} \\ 8 \text{ atoms H} & 3 \text{ atoms C} \\ \text{propane} & \text{propane} \\ 10 \text{ atoms H} & 4 \text{ atoms C} \\ \text{butane} & \text{butane} \end{array} \right) \]
\[
= \left( 7.0 \times 10^{14} \text{ atoms H, } 2.0 \times 10^{14} \text{ atoms C} \right)
\]

Matrix transpose operations

Above, the lists (or vectors) of atoms were stored as rows, but often it is convenient to store them as columns. The operation to take a row to a column (and vice-versa) is called a “transpose”.

\[
M_{HC}^{\top} = \left( \begin{array}{ccc} \text{number of H} & \text{number of C} \\ \text{methane molecule} & \text{methane molecule} \\ \text{number of methanes} & \text{number of C-atoms} \\ \text{propane molecule} & \text{propane molecule} \\ \text{number of propanes} & \text{number of C-atoms} \\ \text{butane molecule} & \text{butane molecule} \\ \text{number of butanes} & \text{number of C-atoms} \end{array} \right)
\]

\[
N_{HC}^{\top} = \left( \begin{array}{c} \text{number of methanes} \\ \text{number of propanes} \\ \text{number of butanes} \end{array} \right) = \left( \begin{array}{c} N_{HCm} \\ N_{HCP} \\ N_{HCI} \end{array} \right)
\]

\[
M_{HC}^{T} N_{HC}^{\top} = N^{T} \left( \begin{array}{ccc} 4 & 8 & 10 \\ 1 & 3 & 4 \end{array} \right) \left( \begin{array}{c} \text{number of methanes} \\ \text{number of propanes} \\ \text{number of butanes} \end{array} \right) = \left( \begin{array}{c} \text{number of H-atoms} \\ \text{number of C-atoms} \end{array} \right)
\]

Matrix Multiplication
The next example supposes that some process produces hydrocarbons and can be modeled with the pressure $P$ and temperature $T$. Suppose (this is an artificial example) that the number of hydrocarbons produced in one millisecond can be related linearly to the pressure and temperature:

$$\begin{align*}
\text{number of methanes} &= \alpha P + \beta T \\
\text{number of propanes} &= \gamma P + \delta T \\
\text{number of butanes} &= \epsilon P + \phi T
\end{align*}$$

(6-24)

or

$$\vec{N}_{HC} = \begin{pmatrix} \alpha & \beta \\ \gamma & \delta \\ \epsilon & \phi \end{pmatrix} \begin{pmatrix} P \\ T \end{pmatrix}$$

(6-25)

Lecture 06  MATHEMATICA® Example 1

Matrices


Here is an example operation that takes us from the processing vector $(P, T)^T$ to the number of hydrogens and carbons.

1: The matrix (Eq. 6-12) is entered as a list of sublists. The sub-lists are the rows of the matrix. The first elements of each row-sublist form the first column; the second elements are the second column and so on.

The Length of a matrix-object gives the number of rows, and the second member of the result of Dimensions gives the number of columns.

All sublists of a matrix must have the same dimensions.

It is good practice to enter a matrix and then display it separately using MatrixForm. Otherwise, there is a risk of defining a symbol as a MatrixForm-object and not as a matrix which was probably the intent.

2: The Transpose function exchanges the rows and columns. If Dimensions[Mat] returns \{r,c\}, then Dimensions[Transpose[Mat]] returns \{c,r\}.

3: Dimensions[PTmatrix] is \{3,2\}.

4: This command will generate an error.

Matrix multiplication in MATHEMATICA® is produced by the "dot" (.) operator—and not the "multiplication" ( * ) operator. For matrix multiplication, $A.B$, the number of columns of $A$ must be equal to the number of rows of $B$.

6: The Transpose "flips" a matrix by producing a new matrix which has the original’s $i^{th}$ row as the new matrix’s $i^{th}$ column (or, equivalently the $j^{th}$ column as the new $j^{th}$ row). In this example, a $3 \times 2$-matrix (PTmatrix) is being left-multiplied by a a $2 \times 3$-matrix. The resulting matrix would map a vector with values $P$ and $T$ to a vector for the rate of production of C and H.
Matrix multiplication is defined by:

$$AB = \sum_i A_{ki}B_{ij}$$  \hspace{1cm} (6-26)

The indices of the matrix defined by the multiplication $AB = C$ are $C_{kj}$.

Matrix Inversion

Sometimes what we wish to know is: “What vector is it ($\vec{x}$), when transformed by some matrix ($A$), that gives us a particular result ($\vec{b} = A\vec{x}$)?”

$$A\vec{x} = \vec{b}$$

$$A^{-1}A\vec{x} = A^{-1}\vec{b}$$

$$\vec{x} = A^{-1}\vec{b}$$  \hspace{1cm} (6-27)

The inverse of a matrix is defined as: something, that when multiplied with the matrix, leaves a product that has no effect on any vector. This special product matrix is called the identity matrix.
Lecture 06  MATHEMATICA® Example 2

Inverting Matrices


Our last example produced a linear operation that answered the question, “given a particular $P$ and $T$, at what rate will $C$ and $H$ be produced?”

To answer the converse question, “If I want a particular rate of production for $C$ and $H$, at what $P$ and $T$ should the process be carried out?”

To invert the question on linear processes, the matrix is inverted.

\[
\begin{align*}
MPT &= \text{Transpose}(AMHCE) \text{ PTmatrix;} \\
MPT^{-1} &= \text{Factor}(\text{Inverse}(MPT)); \\
MPT^{-1} &\equiv \text{MatrixForm}
\end{align*}
\]

1: Inverting a matrix by hand is tedious and prone to error, \textbf{Inverse} does this in \textsc{Mathematica}®. In this example, \textbf{Factor} is called on the result of \textbf{Inverse}. \textbf{Factor} is an example of a \textit{threadable function}—it recursively operates on all members of any argument that is a list-object. Thus, each of the entries in the inverted matrix is factored individually.

2: The \textit{determinant} of a matrix is fundamentally linked to the existence of its inverse. In this example, it is observed that if the \textbf{Det} of a matrix vanishes, then the entries of its inverse are undetermined.

3: The multiplication of a matrix by its inverse should produce the identity matrix (i.e., a matrix with 1 at each diagonal entry, and zero otherwise). That this multiplication gives the identity matrix is not obvious. Unless, …

4: \textbf{Simplify} is called on each of the entries.
Linear Independence: When solutions exist

\[ \begin{pmatrix} a_{11} & a_{12} \\ a_{21} & a_{22} \end{pmatrix} \begin{pmatrix} x \\ y \end{pmatrix} = \begin{pmatrix} b_1 \\ b_2 \end{pmatrix} \quad (6-28) \]

No Solution One Unique Solution Infinitely Many Solutions

Figure 6-1: Geometric interpretation of solutions in two dimensions

Figure 6-2: Geometric interpretation of solutions in three dimensions. From left to right: One solution, No solutions. Infinitely many solutions (two dimensional set) Infinitely many solution (one dimensional set).
Consider liquid water near the freezing point—dipole interactions will tend to make water molecules form clusters such as H₂O and H₄O₂.

This example looks at such a case where the columns are not linearly independent.

1: The mapping from molecules to the number of atoms becomes:

\[
\begin{pmatrix}
2 & 4 \\
1 & 2 \\
\end{pmatrix}
\begin{pmatrix}
N_{\text{H}_2\text{O}} \\
N_{\text{H}_4\text{O}_2} \\
\end{pmatrix}
=\begin{pmatrix}
N_{\text{H}} \\
N_{\text{O}} \\
\end{pmatrix}
\]  

(6.29)

The matrix \( \text{watmat} \) encodes the coefficients in these linear equations.

2–5: The vectors, \( \text{atomvec} \) and \( \text{molvec} \), represent the numbers of each type of atom and each type of molecule.

5–6: These equations are the same as the rows of \( A\vec{x} \) being set to the corresponding entry of \( \vec{b} \) for \( A\vec{x} = \vec{b} \). These are the linear equations given above.

7: This is an attempt (using \texttt{Solve} on the linear equations) to find the number of H₂O- and H₄O₂-molecules, given the number of H- and O-atoms. Of course, it has to fail.

8–9: \texttt{Eliminate} produces a logical equality for each redundancy in a set of equations. In this case, the result expresses the fact that 2 \times (second row) is the same as the (first row).

10: The rank of a matrix, obtained with \texttt{MatrixRank}, gives the number of linearly independent rows.

11: The null space of a matrix, \( A \), is a linearly independent set of vectors \( \vec{x} \), such that \( A\vec{x} \) is the zero-vector; this list can be obtained with \texttt{NullSpace}. The result is equivalent to that obtained with \texttt{Eliminate} in item 9. The \textit{nullity} is the number of vectors in a matrix’s null space. The rank and the nullity must add up to the number of columns of \( A \).